

Download Mole Conversions Using Avogadro's Number Worksheet

Worksheet #13 Using Avogadro's number and Molar Masses (2p.) 2 Answers. 1. All answers have units of grams/mole AgCl 143.32g/mol CaBr₂ 199.88 Li₂O 29.88 C₆H₁₂O₆ 180.18 Fe(NO₃)₃ 241.88 Al₂(SO₄)₃ 342.17 All answers to #2, #3 and #4 should have the chemical formula as part of the unit. 2. a) 152.3g HCN Calculate the number of molecules in 1.058 mole (or gram) of H₂O Calculate the number of atoms in 0.750 mole (or gram) of Fe These problems use the reverse technique of the above. Once again, replacing mole with gram adds one step to the procedure. Here is a graphic of the procedure steps: Mole Practice Name: 23 Avogadro's Number (6.02×10^{23}) was not discovered by him, but named in honor of him. It was Dr. Avogadro's original hypothesis about the volume of gas molecules that led to the development of the mole concept many years later. Use your newfound molar repertoire to complete the following Date. 2. Calculate grams in 3.0000 moles of CO₂ 132.03g 3. Calculate number of moles in 32.0 g of CH₄ 2.00mol 4. Determine mass in grams of 40.0 moles of Na₂CO₃ 4240g 5. Calculate moles in 168.0 g of HgS 0.722mol 6. Calculate moles in 510.0 g of Al₂S₃ 3.396mol 7. How many moles are in 27.00 g of H₂O 1.50mol 8. Determine the mass in grams of Avogadro number of C₁₂H₁₂, Mole Conversions Using Avogadro's Number Worksheet.

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